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The most common measure of concentration used by chemists is molarity (M). Molarity is defined as the number of moles of solute (mol) divided by the total volume (V) of the solution in liters (L). $Molarity = \text{moles of solute per liter of solution}$ ($M = \text{mol} / \text{L}$). Molarity also is called molar concentration.

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11/19: Molarity POGIL (both in-class), No HW 11/20: Molarity and Dilution Practice (answer keys in packet)- complete front AND back of last page for HW. [CLICK HERE](#) for video 11/21: Slushy Lab outside- wear warm clothing and bring gloves 11/22: Creating Solutions Lab- quiz grade for preparing solutions Week 2: 12/2: Acid/Base Cut and Sort ...

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chemists is molarity (M). Molarity is defined as the number of moles of solute (mol) divided by the total volume of the solution in liters (L). Recall the that volume of the solution is the sum of the volume of solvent and the volume of solute. €
Molarity= molesofsolute literofsolution =M= mol L Revised 2017-2018 LCA

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solution is defined as the number of moles of solute per one liter of solution.
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Molarity Practice I 11. How many grams of potassium chloride are needed to make 200 ml, of a 2.5 M solution? (Answer: $2.5\text{M} = 0.5\text{mol} \times 0.2\text{L} = 0.1\text{mol}$)
12. What is the molarity of a 5.00 x 10² ml, solution containing 249 g of calcium iodide? (Answer: 1.69 M)
13. How many grams of calcium chloride are needed to make 0.5 L of a 0.850 M solution? (Answer: 85.0 g)

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